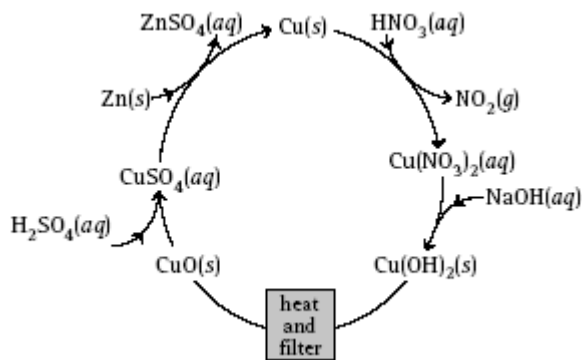


Name _____
Date _____ Period _____

UNIT 1 Chapter 2 Quiz
Multiple Choice Form B

Select the best answer. (1 point each)

1. CaCl_2 is an example of
 - A. a metal.
 - B. a phase.
 - C. a compound.
 - D. an element.
2. Which statement is *not* a description of a chemical reaction?
 - A. Two liquids are mixed in a beaker and a solid forms at the bottom of the beaker.
 - B. A solid is dropped into a beaker of liquid and an explosion occurs.
 - C. A liquid is heated on a Bunsen burner and its temperature rises.
 - D. A solid is dropped into a beaker of liquid and a gas is released.
3. Suppose you pour nitric acid on some copper powder in a test tube. A brown gas is produced and a light blue liquid is all that remains in the test tube. What happened to the copper atoms?
 - A. They are now in the light blue liquid as copper atoms.
 - B. They are now in the light blue liquid as copper ions.
 - C. They escaped as part of a brown gas.
 - D. They were destroyed by the acid.
4. The copper cycle, shown below, is a demonstration of what principle of chemistry?



- A. Most elements are metallic solids.
- B. Elements exist as gases, liquids, or solids.

- C. Mass is conserved in chemical reactions.
- D. Metals are excellent conductors.

For Questions 5–9, use a periodic table.

5. At room temperature, most of the elements in the periodic table are
 - A. gases.
 - B. metalloids.
 - C. solids.
 - D. nonmetals.
6. Which of these describes barium's location on the periodic table?
 - A. Group 3A, Period 2
 - B. Group 2, Period 3A
 - C. Group 2A, Period 6
 - D. Group 6, Period 2A
7. What is the correct ordering of the groups from left to right in the periodic table?
 - A. noble gases, alkali metals, alkaline earth metals, halogens
 - B. alkaline earth metals, alkali metals, halogens, noble gases
 - C. alkali metals, noble gases, alkaline earth metals, halogens
 - D. alkali metals, alkaline earth metals, halogens, noble gases
8. Which pairs of elements would you expect to have the most similar properties?
 - A. arsenic, As, and silicon, Si
 - B. fluorine, F, and chlorine, Cl
 - C. nitrogen, N, and oxygen, O
 - D. hydrogen, H, and helium, He
9. Which of these elements would you expect to have the greatest reactivity?
 - A. boron, B
 - B. gold, Au
 - C. argon, Ar
 - D. barium, Ba
10. What is the correct way to represent a solution of copper nitrate?
 - A. $\text{Cu}(aq)$
 - B. $\text{Cu}(\text{NO}_3)_2(aq)$
 - C. $\text{Cu}(s)$
 - D. $\text{Cu}(\text{NO}_3)_2(s)$